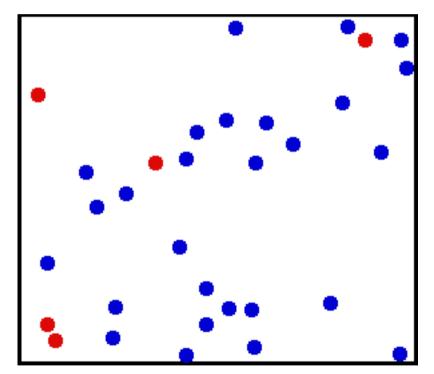
Ludwig Boltzmann 1844-1906



Statistical Mechanics, and kinetic theory of gases, Based on atoms and molecules. Opposed Mach & Oswald: energy is all, no atoms.

Second law of Thermodynamics: Entropy increase -Disorder increases.

$$S=k_B . InW$$



Bolzmann's equation: distribution of molecules position and speed f

$$\left. \frac{\partial f}{\partial t} + v \frac{\partial f}{\partial x} + \frac{F}{m} \frac{\partial f}{\partial v} = \frac{\partial f}{\partial t} \right|_{\text{collision}}$$